



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
General Certificate of Education Ordinary Level

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COMBINED SCIENCE

5129/01

Paper 1 Multiple Choice

October/November 2007

1 hour

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

* 8 5 7 3 0 8 8 2 8 6 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

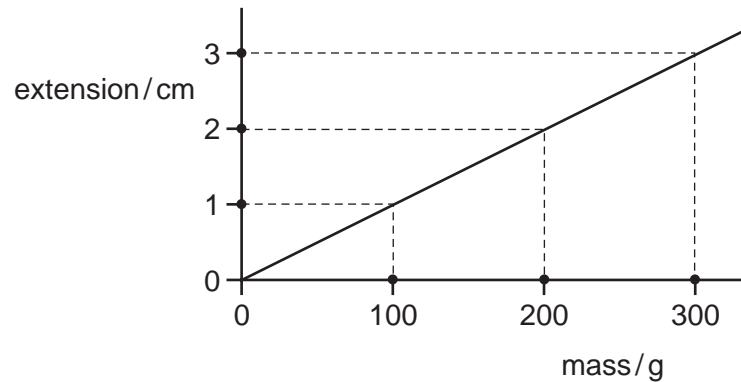
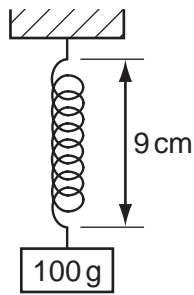
Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 20.

This document consists of **17** printed pages and **3** blank pages.



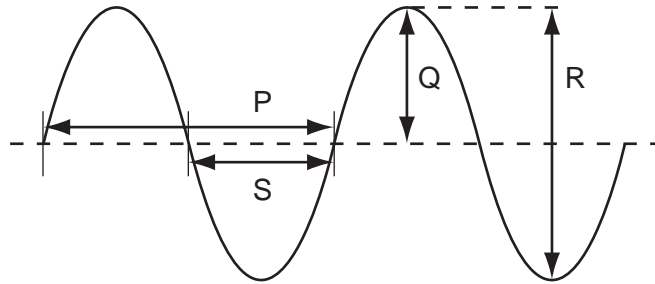
- 1 The diagrams show a spring having a length of 9 cm when loaded with a 100 g mass. The graph shows an extension-mass graph for the spring.



What is the length of the spring after the 100 g mass has been removed?

- A** 7 cm **B** 8 cm **C** 9 cm **D** 10 cm
- 2 Which type of energy is converted to thermal energy when atoms combine?
- A** chemical
B kinetic
C nuclear
D solar
- 3 Equal volumes of four substances are heated at atmospheric pressure.
 The temperature rise is the same for each substance.
 Which substance expands the most?
- A** air
B mercury
C steel
D water

- 4 The diagram shows the surface of the water as a wave passes across a ripple tank.



Which lengths represent the amplitude and wavelength?

| | amplitude | wavelength |
|----------|-----------|------------|
| A | Q | P |
| B | Q | S |
| C | R | P |
| D | R | S |

- 5 A wave has a frequency of 10^4 Hz.

What are the possible values of its velocity and wavelength?

| | velocity in m/s | wavelength in m |
|----------|-----------------|-----------------|
| A | 330 | 0.33 |
| B | 330 | 33 |
| C | 3×10^8 | 30 |
| D | 3×10^8 | 3×10^4 |

- 6 Which type of electromagnetic radiation travels at the highest speed through a vacuum?

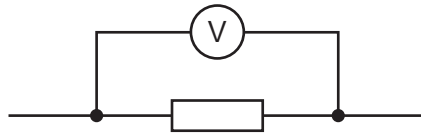
- A** gamma rays
- B** light waves
- C** radio waves
- D** none – all have the same speed

- 7 A lightning flash carries 25 C of charge and lasts for 0.01 s.

What is the current?

- A** 0.0004A
- B** 0.25A
- C** 25A
- D** 2500A

- 8 A voltmeter is connected across a resistor in an electrical circuit.

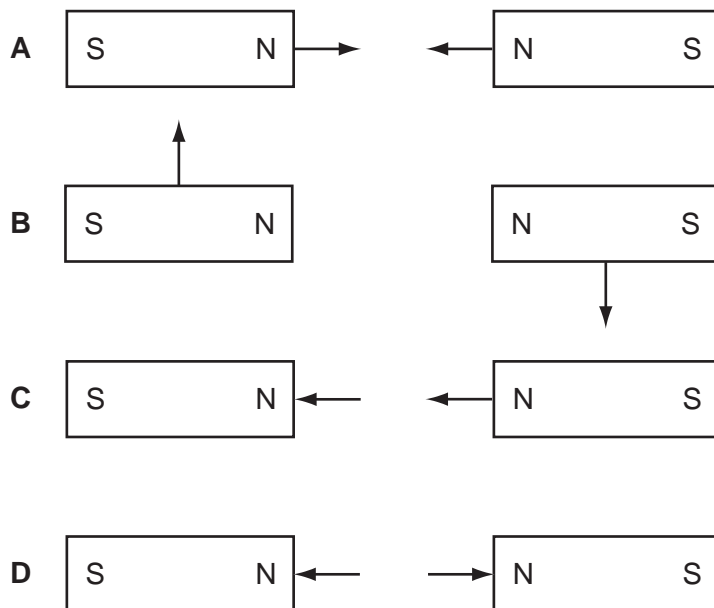


What does the reading on the voltmeter measure?

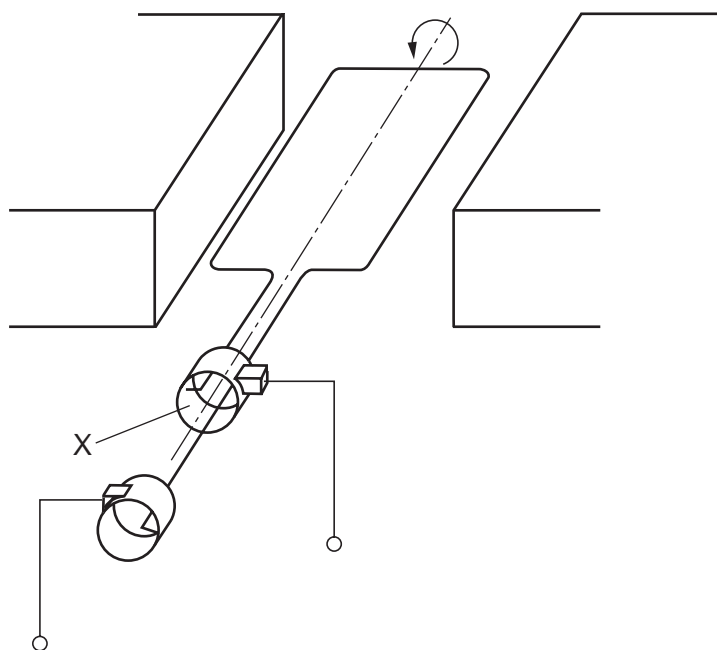
- A** the work done in driving 1 A of current through the resistor
B the work done in driving 1 C of charge through the resistor
C the work done in driving 1 J of energy through the resistor
D the work done in driving 1 W of power through the resistor
- 9 A $1.0\ \Omega$ resistor and a $2.0\ \Omega$ resistor are connected in series across a 12 V d.c. supply.

What is the current in the circuit?

- A** 12 A **B** 6.0 A **C** 4.0 A **D** 0.25 A
- 10 Which diagram shows the correct directions of the magnetic forces on two bar magnets?

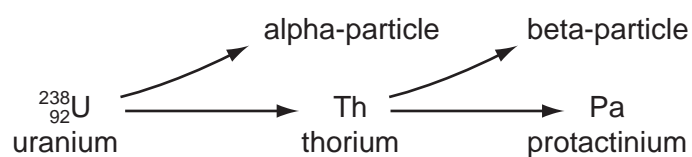


- 11 The diagram shows a simple a.c. generator.



Which name is given to part X?

- A** axle
B carbon brush
C magnet
D slip ring
- 12 The uranium atom ${}_{92}^{238}\text{U}$ emits an alpha-particle to become thorium, which then emits a beta-particle to become protactinium.

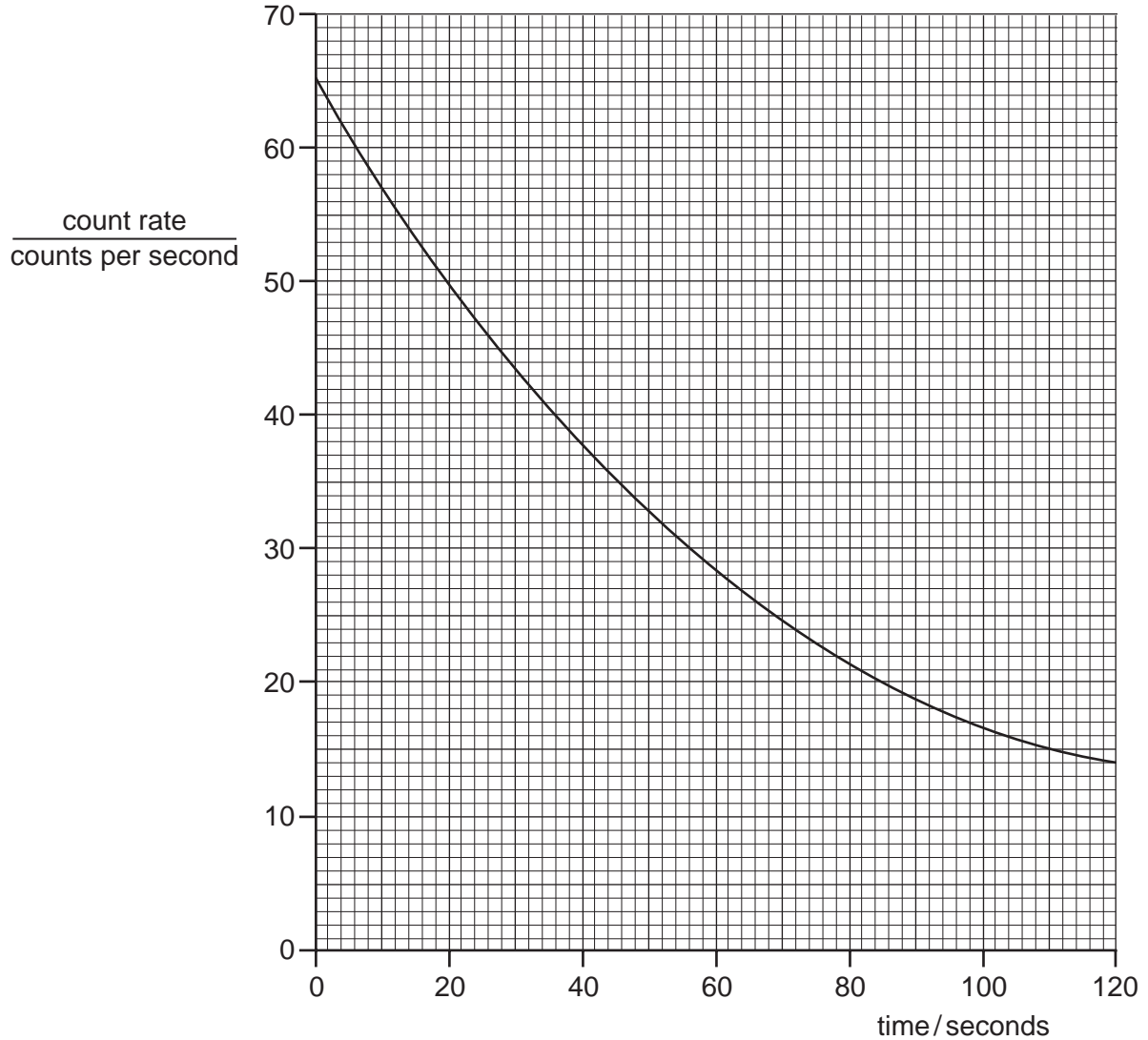


What is the proton number (atomic number) of protactinium?

- A** 89 **B** 90 **C** 91 **D** 95

- 13 Ra decays with a half-life of 1600 s.
Rn decays with a half-life of 52 s.
Po decays with a half-life of 9.1 s.
Pb decays with a half-life of 10.6 h.

The changing count rate for one of these radioactive nuclides is shown in the graph.



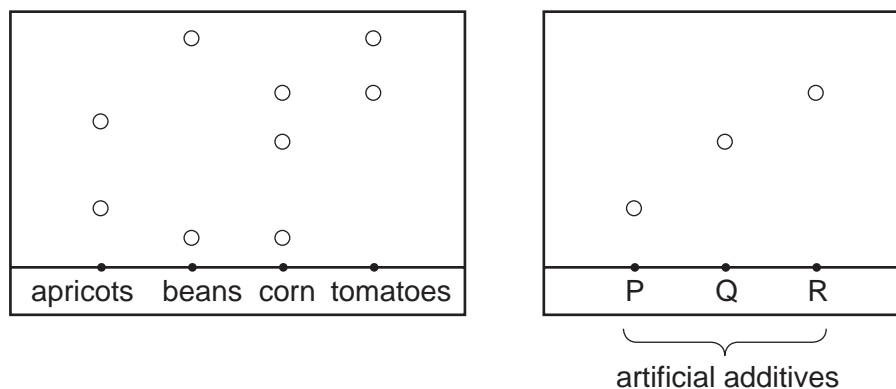
From the half-life shown by the graph, which was the decaying radioactive nuclide?

- A** Ra **B** Rn **C** Po **D** Pb

- 14 Samples of tinned apricots, beans, corn and tomatoes are tested for additives using thin layer chromatography.

The chromatograms are compared with those of three artificial additives, P, Q and R.

The results are as follows.



Which tinned food does **not** contain any artificial additives?

- A apricots
 - B beans
 - C corn
 - D tomatoes
- 15 Element X has proton number 8 and nucleon number 18.

Which particles are present in the X^{2-} ion?

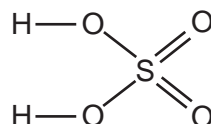
- A 10 electrons, 8 protons, 8 neutrons
 - B 10 electrons, 8 protons, 10 neutrons
 - C 10 electrons, 9 protons, 9 neutrons
 - D 8 electrons, 8 protons, 18 neutrons
- 16 The table gives the electronic structure of four elements.

| element | electronic structure |
|---------|----------------------|
| W | 2.7 |
| X | 2.8.5 |
| Y | 2.8.6 |
| Z | 2.8.8.2 |

Which two elements form an ionic compound?

- A W and X
- B W and Y
- C W and Z
- D X and Y

17 A molecule of sulphuric acid has the structural formula shown.



How many electrons are involved in forming all the covalent bonds in one molecule?

- A** 6 **B** 8 **C** 12 **D** 16

18 The formula of copper(I) oxide is Cu_2O .

How many grams of oxygen are combined with 64 g of copper in this compound?

- A** 8 **B** 16 **C** 32 **D** 64

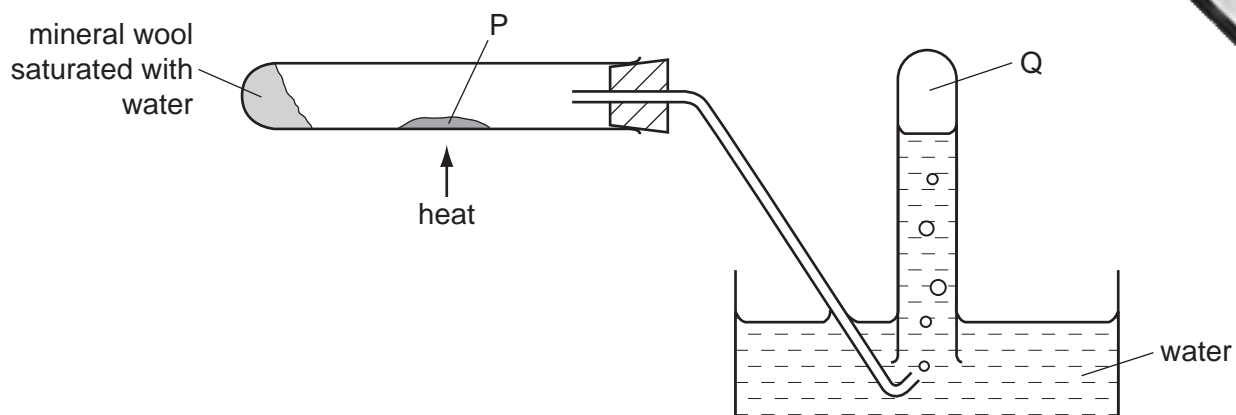
19 Which type of reaction takes place when H^+ ions and OH^- ions react to form water?

- A** condensation
B ionisation
C neutralisation
D precipitation

20 Which statement about the alkali metals is correct?

- A** Their melting points decrease on descending the group.
B Their reactivities decrease on descending the group.
C They form covalent bonds with the halogens.
D They form oxides on reacting with water.

- 21 In the experiment shown in the diagram, steam is passed over a heated solid and a gas is collected.



What are substances P and Q?

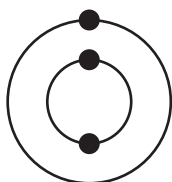
| | P | Q |
|----------|--------|----------|
| A | copper | hydrogen |
| B | lead | oxygen |
| C | silver | oxygen |
| D | zinc | hydrogen |

- 22 The diagrams show the electronic structures of four elements.

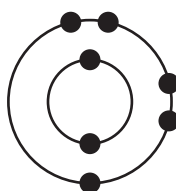
element 1



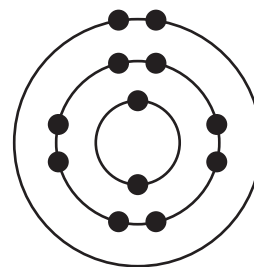
element 2



element 3



element 4



Which two elements are metals?

- A** 1 and 2 **B** 1 and 3 **C** 2 and 4 **D** 3 and 4
- 23 Which substance is added to a blast furnace to remove impurities from iron ore?

- A** carbon
B limestone
C sand
D slag

24 Which pollutant is correctly linked to its source?

| | pollutant | source |
|----------|-----------------|----------------------------|
| A | carbon monoxide | internal combustion engine |
| B | methane | volcanoes |
| C | nitrogen oxide | bacterial decay |
| D | sulphur dioxide | lightning activity |

25 Which statement about the manufacture of ammonia by the Haber Process is correct?

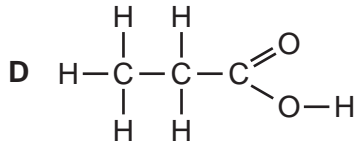
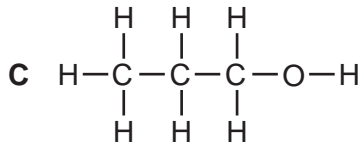
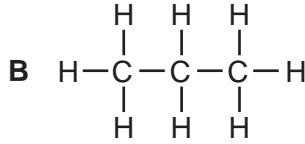
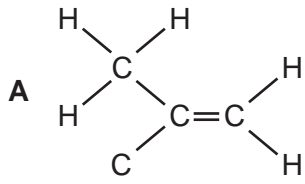
- A** The reactants and product are compounds.
- B** The reactants and product are elements.
- C** The reactants and product are gases.
- D** The reactants are both obtained from the air.

26 Bitumen is obtained from crude oil.

What is it used for?

- A** as fuel for aircraft
- B** as fuel for oil stoves
- C** for making polishes
- D** for making roads

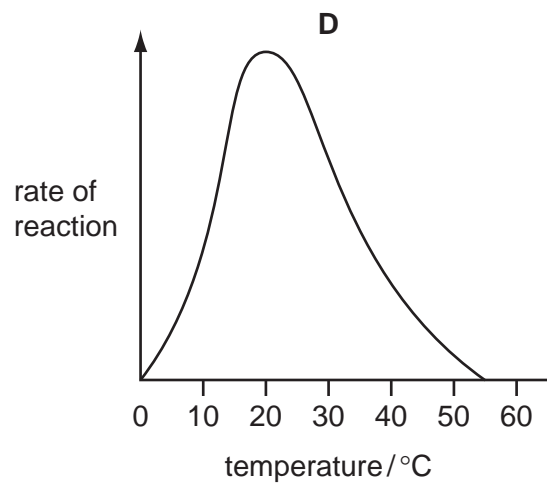
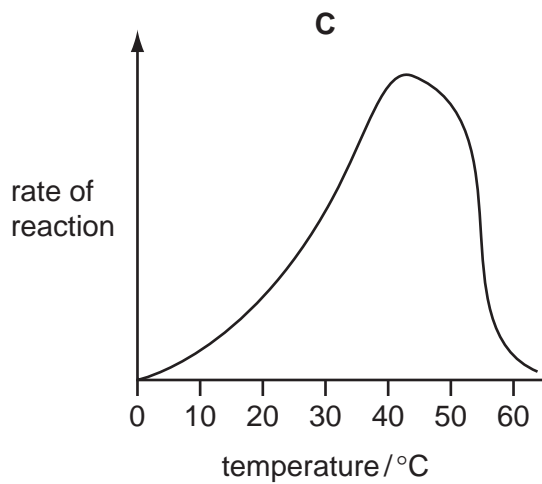
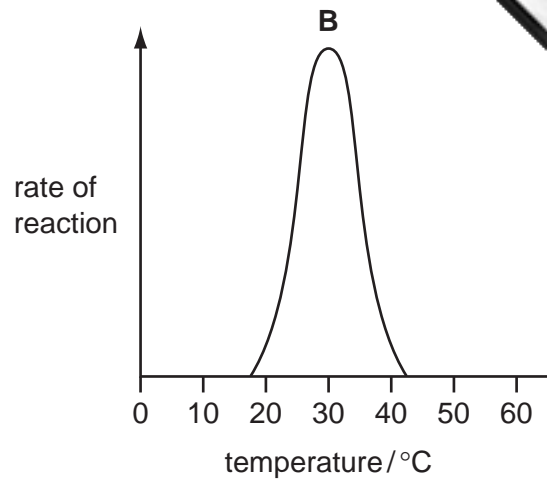
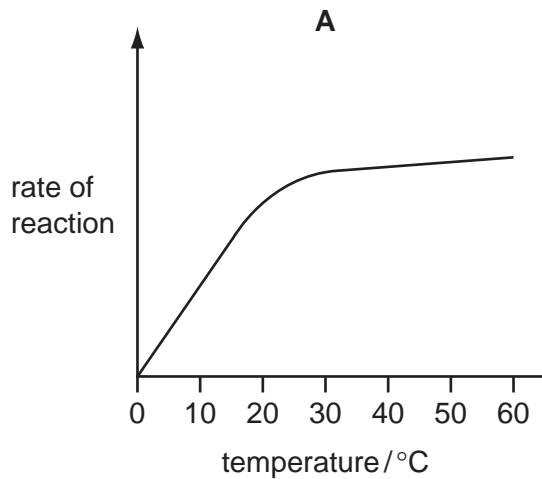
27 Which compound decolourises aqueous bromine?



28 Which cell structure contains the light-absorbing pigments in plants?

- A** chloroplast
- B** cytoplasm
- C** nucleus
- D** vacuole

29 Which graph shows the effect of temperature on enzyme-controlled reactions?



30 How does most carbon dioxide reach the photosynthesising cells of a leaf?

- A through the cuticle
- B through the epidermis
- C through the stomata
- D through the xylem

31 Which part of the alimentary canal is most acidic?

- A colon
- B ileum
- C mouth
- D stomach

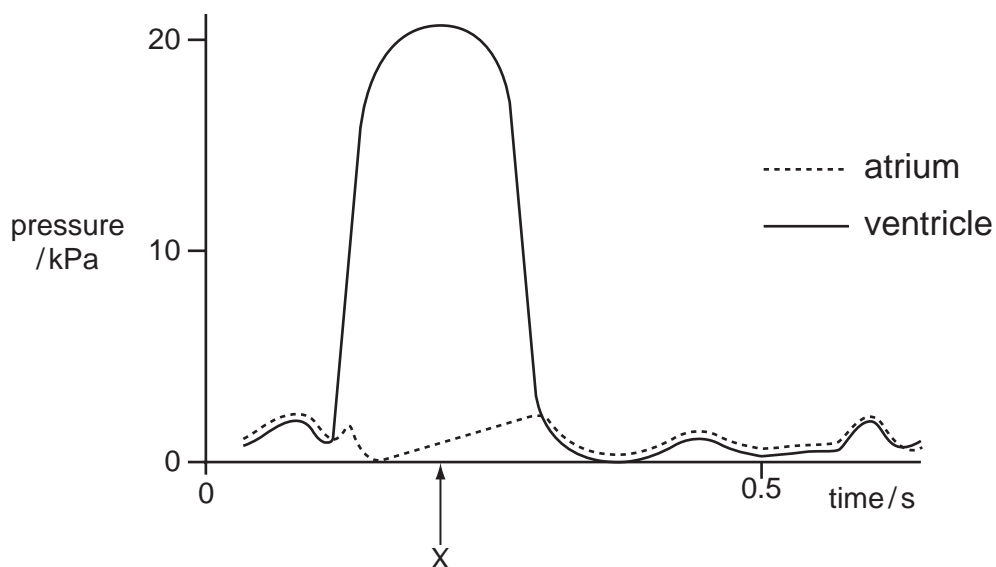
- 32 Four similar leafy shoots are exposed to different conditions. The rates of water uptake and rates of water loss are measured.

The results are shown in the table.

Which shoot is most likely to wilt?

| | water uptake /mm ³ per min | water loss /mm ³ per min |
|----------|--|--|
| A | 10 | 12 |
| B | 10 | 8 |
| C | 5 | 5 |
| D | 5 | 2 |

- 33 The graph shows pressure changes in the left atrium and in the left ventricle during one heartbeat.



What is the state of the valves at time X?

| | bicuspid valve | semi-lunar valve (in aorta) |
|----------|----------------|--------------------------------|
| A | closed | closed |
| B | closed | open |
| C | open | closed |
| D | open | open |

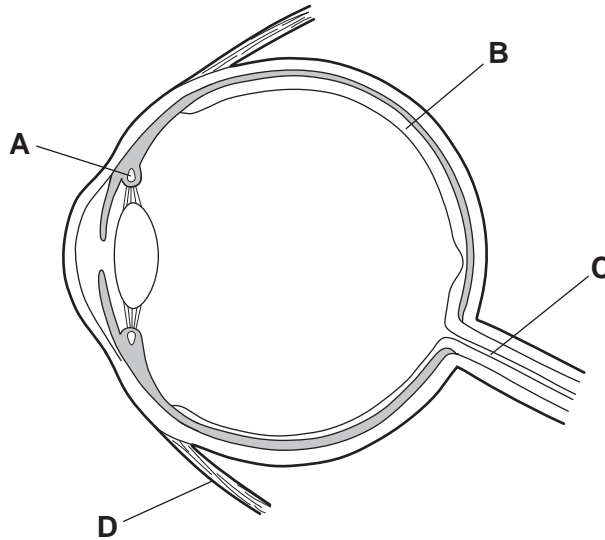
34 The table shows the percentage composition of four samples of air.

Which sample could have been breathed out by a person after vigorous exercise?

| | oxygen | carbon dioxide | water vapour |
|----------|--------|----------------|--------------|
| A | 16 | 0.3 | saturated |
| B | 16 | 4 | saturated |
| C | 21 | 0.03 | trace |
| D | 21 | 3 | trace |

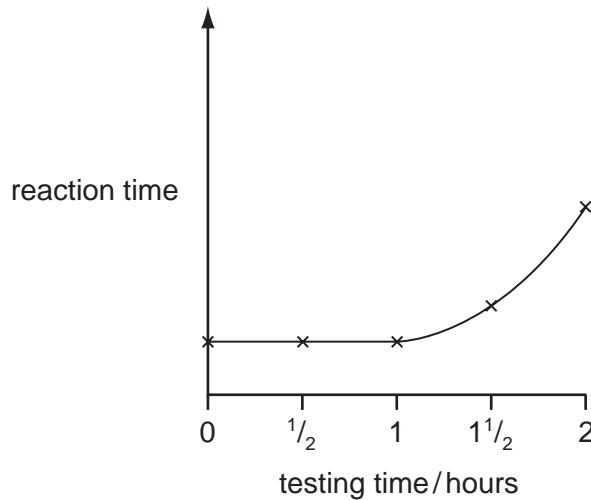
35 The diagram shows a section through an eye.

Which part helps to focus an image on the retina?



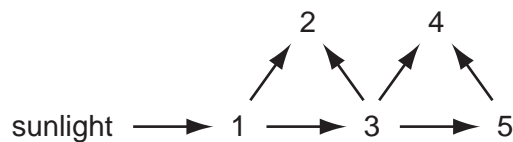
- 36 An experiment was carried out in which the reaction time for a person to respond to a stimulus was measured. Every half hour the person was given an alcoholic drink and the test was repeated.

The results over two hours are shown in the diagram below.



Which deduction can be made from the experiment?

- A Alcoholic drinks make the person react more slowly.
 - B Mental activities are stimulated by small quantities of alcohol.
 - C The alcohol content of the blood rises rapidly after 1 hour.
 - D The person reacts more quickly as a result of practice.
- 37 The diagram shows energy flow in a food web.



Which number represents an organism that eats both plants and animals?

- A 2
 - B 3
 - C 4
 - D 5
- 38 What increases the risk of famine?
- A decreased air pollution
 - B decreased population size
 - C increased carbon dioxide concentration in the air
 - D increased soil erosion

- 39 Which statement is true of asexual reproduction in plants?
- A Insects are needed to transfer pollen.
 - B New plants grow from seeds.
 - C Offspring are genetically identical to their parents.
 - D Two types of gametes are involved.
- 40 What is the path taken by sperm cells during ejaculation from the male reproductive system?
- A sperm duct → testis → urethra
 - B sperm duct → urethra → testis
 - C testis → sperm duct → urethra
 - D testis → urethra → sperm duct

DATA SHEET
The Periodic Table of the Elements

| | | Group | | | | | | | | | | | | | |
|-----------------------------------|------------------------------------|------------------------------------|--|-------------------------------------|--------------------------------------|-------------------------------------|-------------------------------------|--------------------------------------|-------------------------------------|---------------------------------------|---------------------------------------|------------------------------------|--|-------------------------------------|---------------------------------------|
| I | II | III | IV | V | VI | VII | VIII | IX | X | XI | XII | | | | |
| | | 1 H Hydrogen 1 | | | | | | | | | | | | | |
| 7 Li Lithium 3 | 9 Be Beryllium 4 | | | | | | | | | | | | | | |
| 23 Na Sodium 11 | 24 Mg Magnesium 12 | | | | | | | | | | | | | | |
| 39 K Potassium 19 | 40 Ca Calcium 20 | 51 V Vanadium 23 | 52 Cr Chromium 24 | 55 Mn Manganese 25 | 56 Fe Iron 26 | 59 Co Cobalt 27 | 59 Ni Nickel 28 | 64 Cu Copper 29 | 65 Zn Zinc 30 | 70 Ga Gallium 31 | 73 Ge Germanium 32 | 75 As Arsenic 33 | 79 Se Selenium 34 | 80 Br Bromine 35 | 84 Kr Krypton 36 |
| 85 Rb Rubidium 37 | 88 Sr Strontium 38 | 91 Zr Zirconium 40 | 96 Mo Molybdenum 42 | 101 Ru Ruthenium 44 | 103 Rh Rhodium 45 | 106 Pd Palladium 46 | 108 Ag Silver 47 | 112 Cd Cadmium 48 | 115 In Indium 49 | 119 Sn Tin 50 | 122 Sb Antimony 51 | 127 I Iodine 53 | 128 Te Tellurium 52 | 131 Xe Xenon 54 | |
| 133 Cs Caesium 55 | 137 Ba Barium 56 | 181 Ta Tantalum 73 | 184 W Tungsten 74 | 186 Re Rhenium 75 | 190 Os Osmium 76 | 192 Ir Iridium 77 | 195 Pt Platinum 78 | 197 Au Gold 79 | 201 Hg Mercury 80 | 204 Tl Thallium 81 | 207 Pb Lead 82 | 209 Bi Bismuth 83 | 210 Po Polonium 84 | 210 At Astatine 85 | 222 Rn Radon 86 |
| 87 Fr Francium | 88 Ra Radium | 226 Po Polonium | 210 At Astatine | 210 Rn Radon | 210 Ac Actinium | † | | | | | | | | | |
| | | 140 Ce Cerium 58 | 141 Pr Praseodymium 59 | 144 Nd Neodymium 60 | 147 Pm Promethium 61 | 150 Sm Samarium 62 | 152 Eu Europium 63 | 157 Gd Gadolinium 64 | 159 Tb Terbium 65 | 162 Dy Dysprosium 66 | 165 Ho Holmium 67 | 167 Er Erbium 68 | 169 Tm Thulium 69 | 173 Yb Ytterbium 70 | 175 Lu Lutetium 71 |
| | | 232 Th Thorium 90 | 232 Pa Protactinium 91 | 238 U Uranium 92 | 238 Np Neptunium 93 | 238 Pu Plutonium 94 | 238 Am Americium 95 | 238 Cm Curium 96 | 238 Bk Berkelium 97 | 238 Cf Californium 98 | 238 Es Einsteinium 99 | 238 Fm Fermium 100 | 238 Md Mendelevium 101 | 238 No Nobelium 102 | 238 Lr Lawrencium 103 |

* 58-71 Lanthanoid series
† 90-103 Actinoid series

Key

| | | |
|---|----------|---|
| a | X | b |
|---|----------|---|

a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).